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Chemistry 12 Unit 3 - Solubility of Ionic Substances

16. If 25.0 mL of 0.90 M HCl is added to 125.0 mL of water, what is the final [HCl]? **[KEY]**

$$FC = IC \times \frac{IV}{FV}$$
$$= 0.90 M \times \frac{25.0 \text{ mL}}{150.0 \text{ mL}} = 0.15 M$$

Answer: $[HCl] = 0.15 M$

17. Calculate the $[Fe^{3+}]$ in a 0.25 M solution of $Fe_2(SO_4)_3$?

$$0.25 M \xrightarrow{\frac{2}{1}} 0.50 M$$
$$Fe_2(SO_4)_3(s) \rightarrow 2 Fe^{3+} + 3 SO_4^{2-}$$

Answer: $[Fe^{3+}] = 0.50 M$

18. Calculate the $[Na^+]$ in a 0.55 M solution of sodium acetate. (Write the proper formula for sodium acetate first.)

$$0.55 M \xrightarrow{\frac{2}{1}} 1.1 M$$
$$NaCH_3COO(s) \rightarrow Na^+ + CH_3COO^-$$

Answer: $[Na^+] = 0.55 M$

19. Calculate the $[Na^+]$ in a 0.55 M solution of sodium carbonate. (Write the proper formula for sodium carbonate first.)

$$0.55 M \xrightarrow{\frac{2}{1}} 1.1 M$$
$$Na_2CO_3(s) \rightarrow 2 Na^+ + CO_3^{2-}$$

Answer: $[Na^+] = 1.1 M$

20. Calculate the $[Na^+]$ in a 0.55 M solution of sodium phosphate. (Write the proper formula for sodium phosphate first.)

$$0.55 M \xrightarrow{\frac{3}{1}} 1.65 M$$
$$Na_3PO_4(s) \rightarrow 3 Na^+ + PO_4^{3-}$$

Answer: $[Na^+] = 1.7 M$
(2 SD's)

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