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10000 101 Workbook, Chapter 4 - Gas
Multiple choice questions (100 marks) (100 marks)
Show your work questions (10 marks) (10 marks)
Total = 110 marks (110 marks) (110 marks) (110 marks) (110 marks)

1. 1.00 mol of gas at 1.00 atm and 273 K. What is the volume of the gas?
A. 22.4 L B. 22.4 dm³ C. 22.4 m³ D. 22.4 cm³ E. 22.4 m²

2. How many moles of gas are there in 1.00 dm³ at 1.00 atm and 273 K?
A. 0.0446 mol B. 0.0446 mol C. 0.0446 mol D. 0.0446 mol E. 0.0446 mol

3. What is the volume occupied by 1.00 mol of gas at 1.00 atm and 273 K?
A. 22.4 L B. 22.4 dm³ C. 22.4 m³ D. 22.4 cm³ E. 22.4 m²

4. What is the mass of 1.00 mol of gas at 1.00 atm and 273 K?
A. 2.00 g B. 2.00 g C. 2.00 g D. 2.00 g E. 2.00 g

5. A certain amount of gas at 1.00 atm and 273 K is heated to 546 K. What is the volume of the gas?
A. 44.8 L B. 44.8 dm³ C. 44.8 m³ D. 44.8 cm³ E. 44.8 m²

6. Which of the following gases will have the highest density, when all are at the same temperature and pressure?
A. CH₄ B. CO₂ C. O₂ D. N₂ E. H₂

7. Calculate the density of nitrogen, N₂, at 1.00 atm and 273 K.
A. 1.25 g/L B. 1.25 g/dm³ C. 1.25 g/m³ D. 1.25 g/cm³ E. 1.25 g/m²

8. An unknown gas has a density of 1.96 g/L at 1.00 atm and 273 K. What is the molar mass of the gas?
A. 44.0 g/mol B. 44.0 g/dm³ C. 44.0 g/m³ D. 44.0 g/cm³ E. 44.0 g/m²

9. What volume of 1.00 mol of gas at 1.00 atm and 273 K is collected at the same pressure and 546 K? The reaction is 2H₂ + O₂ → 2H₂O.
A. 22.4 L B. 22.4 dm³ C. 22.4 m³ D. 22.4 cm³ E. 22.4 m²

10. A mixture of 1.00 mol of gas and 2.00 mol of gas has a total pressure of 1.00 atm. What are the partial pressures of the two gases?
A. 0.33 atm, 0.67 atm B. 0.33 atm, 0.67 atm C. 0.33 atm, 0.67 atm D. 0.33 atm, 0.67 atm E. 0.33 atm, 0.67 atm

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Chemistry 101 Chapter 4 Answers