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Name _____ Date _____ Class _____

CHAPTER 13 REVIEW
Ions in Aqueous Solutions and Colligative Properties

SECTION 2

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. _____ a. Predict the boiling point of a 0.200 m solution of glucose in water.

_____ b. Predict the boiling point of a 0.200 m solution of potassium iodide in water.

2. A chief ingredient of antifreeze is liquid ethylene glycol, $C_2H_4(OH)_2$. Antifreeze $C_2H_4(OH)_2$ is added to a car radiator that is holding 5.0 kg of water.

_____ a. How many moles of ethylene glycol should be added to the radiator to lower the freezing point of the water from 0°C to -18°C ?

_____ b. How many grams of ethylene glycol does the quantity in part a represent?

_____ c. Ethylene glycol has a density of 1.1 g/mL . How many liters of $C_2H_4(OH)_2$ should be added to the water in the radiator to prevent freezing down to -18°C ?

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