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ANSWER KEY

Chapter 17	Section 1	Section 2	Section 3
1. bubble chamber	1. atomic number	1. isotopes are atoms of the same element that differ only in the number of neutrons in their nuclei.	1. atomic number
2. particle accelerator	2. chemical symbol	2. The isotope boron-10 has one less neutron than boron-11.	2. element
3. Democritus	3. average atomic mass	3. The average atomic mass of an element is the average mass of the mixture of its isotopes.	3. chemical symbol
4. Aristotle	4. groups (or families)	4. The two chlorine isotopes have the same number of protons in their nuclei and the same number of electrons around their nuclei. However, chlorine-37 has two more neutrons in its nucleus than does chlorine-35.	4. average atomic mass
5. electron cloud	5. properties	5. Because these isotopes occur in equal parts in nature, the average atomic mass will be simply an average of the three atomic masses. Therefore, the average atomic mass will be 204 amu.	5. groups (or families)
6. Thomson	6. periods		6. electrons
7. orbits	7. metals		7. properties
8. Dalton	8. transition elements		8. periods
9. atom			9. metals
10. Rutherford			10. transition elements
11. quark			11. nonmetals
12. Chadwick			
13. Dalton's model: it was corrected by adding charged particles inside			
14. Thomson's model: it was corrected by adding a nucleus			
15. Rutherford's model: it was corrected by adding energy levels			
16. Bohr's model: it was corrected by giving the electrons clouds of probability instead of little orbit tracks to run on			

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